

## **METHODS AND EQUIPMENT FOR REMOVING STAINS FROM FABRICS**

This application is a continuation-in-part application of U.S. Patent Application Serial No. 10/373,787, filed February 27, 2003, entitled "METHODS AND EQUIPMENT FOR REMOVING STAINS FROM FABRICS," which is incorporated herein by reference.

### **TECHNICAL FIELD**

This invention relates to methods and kits useful for removing stains, such as menstrual fluid or underarm perspiration stains, from clothes and other soft fabric articles. This invention also relates to methods for reducing the damaging effect of hypochlorite-containing solution on cotton and other soft fabrics.

### **BACKGROUND**

Menstrual fluid, a composition of blood and endometrial cells, is difficult to remove from cotton panties once it has stained the fabric. Ultra Clorox® Regular Bleach is one of the leading household products used for the purpose of cleaning white cotton panties of menstrual fluid stain. Ultra Clorox® Regular Bleach is a designated trademark of the Clorox Company. A typical, undiluted Ultra Clorox Regular Bleach solution contains 6-7.35 wt % of sodium hypochlorite and less than 0.2 wt % of sodium hydroxide. The pH of the undiluted Clorox Bleach solution is around 11.4. Like other chlorine-releasing bleaches, Clorox Bleach, even diluted, will disintegrate the fabric. Moreover, even after lengthy soaking, a dark residue stain may still remain on the cotton fabric, even with scrubbing. Vigorous scrubbing accelerates deterioration of the bleach-weakened cotton fibers which, again, leads to damaged panties, and expense and frustration. Some household products, such as hydrogen peroxide, produce free oxygen to dislodge menstrual fluid discharge from cotton fabric but this process may be effective only when the discharge is fresh and minimal fluid penetration of the fabric has occurred.

Perspiration stain in the underarm areas of white cotton fabric shirts and blouses is also difficult to remove, even for professionals in the garment laundry and cleaner business. Often the stain is not completely removed.

There is a clamor among women around the world for a process that they can use to remove fresh, set-in or old menstrual fluid or perspiration stain from white cotton fabric, and that can do so easily, rapidly, with little or no scrubbing, and with no damage to the cotton fabric.

## SUMMARY OF THE INVENTION

One object of the present invention is to provide methods for reducing the damaging effect of hypochlorite-containing solutions on soft fabrics. The fabrics can be made of cotton, cotton/polyester, or other materials. The fabrics may be, for example, in white.

In accordance with one aspect of the present invention, the method comprises the steps of (a) modifying a hypochlorite-containing solution by adding an alkali metal hydroxide to the solution, such that the weight concentration ratio of the alkali metal hydroxide over the hypochlorite salt in the modified solution is no less than 1:12.5; and (b) contacting the modified solution with a stain on a soft fabric article for at least one minute to remove the stain. In certain cases, the contact with the stain can last for at least 5, 10, 15, 30, 60 minutes or longer before the stain is cleaned.

The stain can be any type of hard-to-remove stains, such as fresh, set-in or old menstrual fluid or underarm perspiration stains. Other examples of hard-to-remove stains include, but are not limited to, those caused by wine, grass, urine, feces, and certain types of ink.

In a preferred embodiment, the alkali metal hydroxide is sodium hydroxide, and the hypochlorite salt is sodium hypochlorite. The weight concentration ratio of sodium hydroxide over sodium hypochlorite in the modified solution can be no less than 1:10, 1:5, 1:2 or 1:1. A higher sodium hydroxide/sodium hypochlorite ratio can also be used.

In one embodiment, the modified solution includes at least 0.2, 0.3, 0.5, 1, 2, 3 or higher weight percent of sodium hydroxide. For instance, the weight percentage of sodium hydroxide can range from about 0.5% to about 3%.

1 In another embodiment, the modified solution includes about 2.5 weight percent of  
2 sodium hypochlorite and 0.5 to 1.25 weight percent of sodium hydroxide. In yet another  
3 embodiment, the modified solution includes about 6 weight percent of sodium hypochlorite  
4 and 1.2 to 3 weight percent of sodium hydroxide

5 In accordance with another aspect of the present invention, the method for reducing  
6 the damaging effect of a hypochlorite salt-containing solution comprises the steps of (a)  
7 modifying the solution by adding an alkali metal hydroxide to the solution, such that the pH  
8 of the modified solution is at least 11.8; and (b) contacting the modified solution with a stain  
9 on a soft fabric article for at least one minute to remove the stain. The fabric article may be,  
10 for example, in white.

11 The pH of the modified solution can be at least 12, 12.5 or 13. In one embodiment,  
12 the pH of the modified solution is about 13.

13 In a preferred embodiment, the alkali metal hydroxide is sodium hydroxide, and the  
14 hypochlorite salt is sodium hypochlorite. The weight percentage of sodium hypochlorite in  
15 the modified solution can be at least 0.1%, 0.5%, 1%, 2%, 3%, 4%, 5%, 6% or more.

16 In one embodiment, the modified solution is a modified form of Ultra Clorox Bleach  
17 Regular. Ultra Clorox Bleach Regular typically contains about 6 weight percent of sodium  
18 hypochlorite and less than 0.2 weight percent of sodium hydroxide. To make the modified  
19 form, an additional amount of sodium hydroxide is added.

20 Another object of the present invention is to provide methods and kits useful for  
21 removing hard-to-remove stains from soft fabric articles. The soft fabric articles can be,  
22 for example, panties, shirts, blouses, pants, jeans, trousers or other soft fabric articles.  
23 The removal preferably is accomplished with little or no scrubbing of the fabrics.

24 In accordance with one aspect of the present invention, the method includes the  
25 steps of (a) providing a cleaning composition which contains an effective amount of a  
26 metallic salt of hypochlorous acid and at least 0.2 weight percent of an alkali metal  
27 hydroxide; and (b) contacting the cleaning composition with a stain on a soft fabric  
28 article for at least one minute.

29 In one embodiment, the metallic salt of hypochlorous acid is sodium hypochlorite,  
30 and the alkali metal hydroxide is sodium hydroxide. The cleaning composition can  
31 include, for example, at least 0.3 weight percent of sodium hydroxide. Preferably, the

cleaning composition contains about 0.5 to about 3 weight percent of sodium hydroxide. In one embodiment, the weight concentration ratio of sodium hypochlorite over sodium hydroxide is about 2:1.

The stain to be removed can be menstrual fluid or underarm perspiration stain. The contact between the cleaning composition and the stain can last at least five, fifteen, thirty minutes, or longer, with no damage to the soft fabric article.

In accordance with another aspect of the present invention, the method includes the steps of (a) providing a cleaning composition which contains an effective amount of a metallic salt of hypochlorous acid and has a pH of at least 11.8; and (b) contacting the cleaning composition with a stain on a soft fabric article for at least one minute. The metallic salt of hypochlorous acid preferably is sodium hypochlorite.

In one embodiment, the cleaning composition contains at least 0.3 weight percent of sodium hydroxide. In another embodiment, the cleaning composition contains about 0.5 to about 3 weight percent of sodium hydroxide.

The pH of the cleaning composition can be, for example, at least 12, 12.5, or 13. The cleaning composition can contact with the stain on the soft fabric article for at least five, fifteen, thirty minutes, or longer, with no damage to the fabric article.

In accordance with yet another aspect of the present invention, a kit is provided that is useful for removing stains from clothes or other soft fabrics. The kit includes a cleaning composition which contains an effective amount of a metallic salt of hypochlorous acid and at least 0.2 weight percent of an alkali metal hydroxide. The kit also has an instruction indicating that the cleaning composition contained therein can be used for removing stains from soft fabric articles.

The metallic salt of hypochlorous acid preferably is sodium hypochlorite, and the alkali metal hydroxide preferably is sodium hydroxide. In one embodiment, the cleaning composition comprises about 0.5 to about 3 weight percent of sodium hydroxide. In one embodiment, the weight concentration ratio of sodium hypochlorite over sodium hydroxide is about 2:1. In another embodiment, the kit includes a spray bottle capable of spraying the cleaning composition onto the soft fabric article.

In accordance with still yet another aspect of the present invention, the kit includes (a) a cleaning composition which contains an effective amount of a metallic salt

1 of hypochlorous acid and which has a pH of at least 11.8; and (b) an instruction for  
2 removing stains from soft fabric articles employing the cleaning composition. The  
3 metallic salt of hypochlorous acid preferably is sodium hypochlorite. In one  
4 embodiment, the cleaning composition includes 0.5-3 weight percent of sodium  
5 hydroxide.

6 In accordance with a further aspect of the present invention, the kit contains (a) a  
7 first compartment which includes a sodium hypochlorite solution which preferably has a  
8 pH of between 11 and 13; (b) a second compartment which includes a sodium hydroxide  
9 solution; and (c) an instruction for removing the stain from the soft fabric article  
10 employing the kit.

11 Other features, objects, and advantages of the present invention are apparent in  
12 the detailed description that follows. It should be understood, however, that the detailed  
13 description, while indicating preferred embodiments of the present invention, is given by  
14 way of illustration only, not limitation. Various changes and modifications within the  
15 scope of the invention will become apparent to those skilled in the art from the detailed  
16 description.

## 17 18 **DETAILED DESCRIPTION**

19 The present invention is based on the surprising discovery that a cleaning  
20 composition which contains a metallic salt of hypochlorous acid and an appropriate  
21 amount of alkali metal hydroxide is effective for removing hard-to-remove stains from  
22 clothes and other soft fabric articles. The metallic salt of hypochlorous acid preferably is  
23 sodium hypochlorite. The alkali metal hydroxide preferably is sodium hydroxide. Other  
24 hypochlorous salts and/or alkali metal hydroxides can also be used in the present  
25 invention.

26 Sodium hypochlorite ( $\text{NaOCl}$ ) dissolves in water to sodium and hypochlorite ions.  
27 The hypochlorite ion is a strong oxidant which can react with numerous materials. The  
28 stability of the sodium hypochlorite solution is affected by the pH of the solution. It has  
29 been reported that sodium hypochlorite is the most stable when the pH of the solution is  
30 between 11 to 13. Such a high pH can be created by adding excess alkali metal  
31 hydroxide, such as sodium hydroxide, to the sodium hypochlorite solution.

1       The decomposition rate of the hypochlorite ion increases when the pH of the  
2 solution falls below 11. This is because of the rapid acid catalyzed decomposition  
3 pathway of the hypochlorite ion. The rate of decomposition also increases when the pH  
4 of the solution is over 13. This is due to the increase in the ionic strength of the solution  
5 caused by the increased level of excess alkali metal hydroxide added to the solution. The  
6 present invention finds that even with a high ionic strength, the sodium  
7 hypochlorite/sodium hydroxide solution is still effective for removing menstrual fluid,  
8 underarm perspiration and other hard-to-remove stains from soft fabric articles. In  
9 addition, the addition of appropriate amounts of alkali metal hydroxide to a hypochlorite  
10 solution retards the damaging effect of the hypochlorite solution on soft fabric (such as  
11 cotton fabric).

12       The concentration of sodium hypochlorite in the cleaning composition of the  
13 present invention preferably is at least 0.1% by weight, based on the total weight of the  
14 cleaning composition. For instance, the concentration of sodium hypochlorite can be at  
15 least 0.5, 1, 2, 3, 4, 5, 6, 7 or 8% by weight. In one embodiment, the concentration of  
16 sodium hypochlorite ranges from 0.1 to 10% by weight. In another embodiment, the  
17 concentration of sodium hypochlorite is about 0.5 to 5% by weight. In yet another  
18 embodiment, the concentration of sodium hypochlorite is about 1 to 2.5% by weight. In  
19 still another embodiment, the concentration of sodium hypochlorite is about 1.5 to 2% by  
20 weight.

21       The concentration of sodium hydroxide in the cleaning composition preferably is  
22 at least 0.2% by weight, based on the total weight of the cleaning composition. For  
23 instance, the concentration of sodium hydroxide can be at least about 0.3, 0.4, 0.5, 1, 1.5,  
24 2, 2.5, 3, 4 or 5% by weight. In one embodiment, the concentration of sodium hydroxide  
25 ranges from about 0.5 to about 3% by weight. In another embodiment, the concentration  
26 of sodium hydroxide ranges from about 1 to 2% by weight. It is generally known that an  
27 appropriate amount of alkali metal hydroxide (such as sodium hydroxide) increases the  
28 stability of sodium hypochlorite in the cleaning composition. Without limiting the  
29 present invention to any particular mechanism, Applicant has found that alkali metal  
30 hydroxide (such as sodium hydroxide) adds significantly to the cleaning power of sodium  
31 hypochlorite to remove stains, such as menstrual fluid or underarm perspiration stains,

1 from clothes and other soft fabric articles while significantly increasing the compatibility  
2 of sodium hypochlorite with soft fabric, such as cotton fabric, thereby preventing sodium  
3 hypochlorite from damaging the fabric.

4 The weight concentration ratio of sodium hydroxide over sodium hypochlorite  
5 may vary substantially without affecting the stain-removing power of the cleaning  
6 composition. Preferably, the weight concentration ratio of sodium hydroxide over  
7 sodium hypochlorite is no less than 1:12.5. For instance, the weight concentration ratio  
8 of sodium hydroxide over sodium hypochlorite can be no less than 1:10, 1:5, 1:2.5 or 1:1.  
9 In one embodiment, the weight concentration ratio of sodium hydroxide over sodium  
10 hypochlorite can range from about 1:5 to about 5:1. In another embodiment, the weight  
11 concentration ratio of sodium hydroxide over sodium hypochlorite is about 1:3 to about  
12 1:1. For instance, the weight concentration ratio of sodium hydroxide over sodium  
13 hypochlorite can be about 1:2.

14 In one embodiment, the cleaning composition includes about 6 weight percent of  
15 sodium hypochlorite and 1.2 to 3 weight percent of sodium hydroxide. In another  
16 embodiment, the cleaning composition includes about 2.5 weight percent of sodium  
17 hypochlorite and 0.5 to 1.25 weight percent of sodium hydroxide. The cleaning  
18 composition of the present invention can be a modified form of regular Clorox Bleach or  
19 Ultra Clorox Bleach with additional sodium hydroxide.

20 The pH of the cleaning composition preferably is at least about 11.8. For  
21 instance, the pH of the cleaning composition can be at least 12, 12.5 or 13. In one  
22 embodiment, the pH of the cleaning composition is about 13.

23 Other ingredients or additives can be added in the cleaning composition. These  
24 ingredients or additives include, for example, chelating agents, phosphorous-containing  
25 salts, surfactants, or abrasive agents. These ingredients or additives, however, are not  
26 necessary for the stain-removing function of the cleaning composition. In one  
27 embodiment, the cleaning composition is free of chelating agents, phosphorous-  
28 containing salts, surfactants, and abrasive agents.

29 In one embodiment, Tilex Instant Mildew Stain Remover<sup>®</sup>, Scrub Free Mildew  
30 Stain Remover<sup>®</sup>, or other off-the-shelf hard-surface cleaners, all of which are marketed  
31 and targeted exclusively as such, are used for removing menstrual fluid, underarm

1 perspiration and other hard-to-remove stains from soft fabrics. Tilex Instant Mildew  
2 Stain Remover and Scrub Free Mildew Stain Remover are designated trademarks of the  
3 Clorox Company and Church & Dwight Company, Inc., respectively. The product labels  
4 and/or use instructions clearly and distinctively warn against using these commercial  
5 cleaners on clothes or soft fabrics, leading the users away from such usages, believing the  
6 compositions are too caustic for clothes or other soft fabrics. Tilex Instant Mildew Stain  
7 Remover contains about 1-5 wt % sodium hypochlorite and about 0.5-2 wt % sodium  
8 hydroxide. The pH of Tilex Instant Mildew Stain Remover is about 12.4 - 12.8. Scrub  
9 Free Mildew Stain Remover contains about 2.3% sodium hypochlorite and less than 1%  
10 sodium hydroxide. The pH of Scrub Free Mildew Stain Remover is about 11.8-12.2.  
11 Other commercial available cleaners that can be used in the present invention include, but  
12 are not limited to, Scrubbing Bubbles Mildew Stain Remover<sup>®</sup> and Lysol Mildew  
13 Remover<sup>®</sup>. Scrubbing Bubbles Mildew Stain Remover and Lysol Mildew Remover are  
14 designated trademarks of SC Johnson and Reckitt Benckiser Inc., respectively.

15 The cleaning composition of the present invention can be stored in a container,  
16 such as a spray bottle, prior to use. Preferably, the container has an instruction indicating  
17 that the enclosed cleaning composition can be used for removing stains, such as  
18 menstrual fluid or perspiration stains, from soft fabric articles.

19 Sodium hypochlorite and sodium hydroxide can be separately stored prior to use.  
20 For instance, they can be stored in two separate compartments of a container. The first  
21 compartment encloses a sodium hypochlorite solution which preferably has a pH of  
22 between 11 and 13. The second compartment encloses a concentrated sodium hydroxide  
23 solution. The two solutions are mixed together upon use. An exemplary device suitable  
24 for this purpose is illustrated in U.S. Patent No. 6,398,077, which is incorporated herein  
25 by reference.

26 Soft fabric articles suitable for the present invention can be made of a variety of  
27 materials, such as cotton or cotton/polyester. The fabric articles preferably are in white  
28 color. Examples of soft fabric articles suitable for the present invention include, but are  
29 not limited to, panties, shirts, blouses, pants, jeans, trousers, and other wear and bed  
30 products.



1           The stains to be removed can be menstrual fluid stains or underarm perspiration  
2 stains. Other hard-to-remove stains, such as wine, grass, urine, feces, or ink stains, can  
3 also be removed using the present invention. The contact between the cleaning solution  
4 and the stain may last for at least one minute before the stain is removed. In one  
5 embodiment, the contact between the cleaning solution and the stain lasts for at least 5,  
6 10, 15, 30, 60 or more minutes before the stain is removed.

7           In accordance with one aspect of the present invention, the soft fabric article that  
8 is to be destained is first soaked in cold water until the stain areas are thoroughly  
9 saturated with water. The fabric article can be swirled around in the water to dislodge as  
10 much stain as possible. For articles heavily soiled with stains, the water may be changed  
11 to repeat the soaking and swirling step.

12           The fabric article is then squeezed to remove excess water. White cotton articles  
13 heavily stained with menstrual fluid may be tinted slightly pink after this step. The  
14 stained areas are arranged for maximal exposure in preparation for the spray with the  
15 cleaning composition. Suitable cleaning compositions used in the present invention  
16 include commercial hard-surface cleaners such as Scrubbing Bubbles Mildew Stain  
17 Remover, Tilex Mildew Remover, Lysol Mildew Remover and Scrub Free Mildew Stain  
18 Remover.

19           The cleaning composition can be sprayed on the stain areas, or the entire article if  
20 necessary. After spraying, the stain areas can be compressed and confined into a small  
21 container to saturate and soak the stain areas or the entire article in the cleaner. In one  
22 instance, two pairs of panties can fit entirely into a pint-sized plastic container.

23           The stained areas are soaked with the cleaning composition until the stain has  
24 been removed. This may require about one to five minutes for removing fresh menstrual  
25 fluid stain, and about thirty minutes to two hours for removing old underarm  
26 perspiration stain. The fabric article can be subsequently inspected for any remaining  
27 stain. If necessary, spot spray can be applied again to remove the remaining stain.

28           After all stain has been removed, the fabric article is thoroughly rinsed in cold  
29 water before being put through the detergent wash/rinse and dry cycle, particularly if the  
30 fabric article is combined with non-colorfast clothing in the wash. Also, this assures that  
31 all sodium hydroxide has been removed from the fabric article before it is worn next to

1 the skin. According to the present invention, menstrual fluid stains or underarm  
2 perspiration stains may be removed from a soft fabric article with little or no scrubbing of  
3 the article.

4 For in-place removal of small menstrual fluid stain spots from white sheets, an  
5 absorbent white toweling may be located underneath the spots. A small amount of spray  
6 is applied and confined to the spotted areas. After stain is gone, the treated areas may be  
7 dampened with a wet cloth to remove the spray product and then allowed to dry.

8 The treated fabric article preferably is not combined with non-colorfast clothing  
9 without first rinsing the treated article thoroughly in cold water. After the stain is  
10 removed, the fabric article preferably is not soaked with the cleaning composition any  
11 longer than necessary.

12 It should be understood that the above-described embodiments and the following  
13 examples are given by way of illustration, not limitation. Various changes and  
14 modifications within the scope of the present invention will become apparent to those  
15 skilled in the art from the present description.

#### 16 17 EXAMPLE I

#### 18 Comparison of Scrubbing Bubbles Mildew Stain Remover, Tilex Mildew Remover, 19 Lysol Mildew Remover and Scrub Free Mildew Stain Remover to Clorox Bleach for the 20 Removal of Menstrual Fluid Stains and Underarm Perspiration Stains

21 Tests reported below show that white cotton fibers have a greater tolerance for  
22 Scrubbing Bubbles Mildew Stain Remover, Tilex Mildew Remover and Lysol Mildew  
23 Remover than for bleaching products like Clorox Bleach. In addition, the spray  
24 application and rapid removal of menstrual fluid stain and underarm perspiration stain  
25 associated with Scrubbing Bubbles Mildew Stain Remover, Tilex Mildew Remover,  
26 Lysol Mildew Remover and Scrub Free Mildew Stain Remover, versus the long  
27 immersed soaking process typical of products currently being used for the same purpose,  
28 indicate that the mildew removers can be used with greater safety on white cotton fabric.

29 Observed was the experimental testing of five common household products; (a)  
30 dilute Clorox Bleach (sodium hypochlorite, 2.4%), (b) Tilex Mildew Remover (sodium  
31 hypochlorite, 2.4%), (c) Lysol Mildew Remover ( sodium hypochlorite, 2.0%), (d)

1 Scrubbing Bubbles Mildew Stain Remover, and (e) Scrub Free Mildew Stain Remover  
2 for the removal of fresh menstrual fluid stain from white cotton (100%) panties. Each of  
3 the four mildew remover products was liberally sprayed on a designated one of four  
4 panty articles, resulting in excellent removal of the stains from each panty article in less  
5 than 1 minute. A pair of similarly soiled panties was soaked in Clorox Bleach for an hour  
6 but the test was terminated, with stain still remaining, because of concern for Clorox  
7 Bleach damage to the panties. The remaining stain was quickly and successfully treated  
8 with one of the mildew remover products.

9 Two additional white cotton panties with set-in menstrual fluid stain were treated  
10 with the product known as Shout<sup>®</sup> (label instructs the user to soak clothing with set-in  
11 stains in Shout overnight or longer) but Shout failed to remove the stains which,  
12 subsequently, resisted several wash and dry cycles. Shout<sup>®</sup> is a designated trademark of  
13 S.C. Johnson. These set-in residue stains were sprayed with Tilex Mildew Remover. For  
14 the first pair panties, a single spray application of Tilex Mildew Remover completely  
15 removed the set-in residue stain in 7 minutes. For the second pair of panties, four spray  
16 applications (a total of 15 squirts) and 30 minutes were required for 95%-99% removal of  
17 the set-in residue stain. At least a dozen successful tests followed, using the mildew  
18 removers on white cotton panties stained with menstrual fluid.

19 Experimental observations of Clorox Bleach, Scrubbing Bubbles Mildew Stain  
20 Remover, Tilex Mildew Remover, Lysol Mildew Remover, and Scrub Free Mildew Stain  
21 Remover were conducted to study the extent of physical damage to cotton cloth that may  
22 be caused by these products. An approximate 10 cm<sup>2</sup> patch of white 100% cotton cloth  
23 (panty crotch thickness) was immersed in 10 ml of the Clorox product. Likewise, similar  
24 patches were immersed in 10 ml each of the mildew removal products. Within four  
25 hours, the patch soaked in Clorox was shredded. After 5 to 6 hours, the patch soaked in  
26 Scrub Free Mildew Stain Remover began to shred. After eight hours, the patches soaked  
27 in the remaining three mildew removal products were taken out of their solutions, dried,  
28 stretched and found to be intact.

29 Tests were conducted to determine the effectiveness of Scrubbing Bubbles  
30 Mildew Stain Remover, Tilex Mildew Remover, Lysol Mildew Remover, and Scrub Free  
31 Mildew Stain Remover on perspiration stain, one of the most difficult stains to remove

1 from the underarms of shirts and blouses. A white shirt, 65% polyester and 35% cotton,  
2 was the test material. A years-old yellowish-brown perspiration stain was embedded in  
3 the seams and fabric of the underarm areas of the sleeves, having stubbornly resisted  
4 many wash and dry cycles. The stained areas of the sleeves were immersed in cold water  
5 for 30 minutes. Then the stained areas were sprayed liberally with Tilex Mildew  
6 Remover and stuffed into a pint-sized plastic container, and allowed to stand for 1 hour.  
7 A barely visible yellowish-brown coloration on portions of the seams still remained but  
8 this disappeared completely after a brief scrubbing between the hands in the spray  
9 product that was left in the fabric. Then the shirt was put through a normal wash and dry  
10 cycle. Six undershirts with old, heavily baked-in underarm perspiration stains, assumed  
11 impossible to remove, were successfully processed: one by Scrubbing Bubbles Mildew  
12 Stain Remover, three by Tilex Mildew Remover, one by Lysol Mildew Remover, and one  
13 by Scrub Free Mildew Stain Remover. Scrubbing Bubbles Mildew Stain Remover  
14 required 40 minutes to remove completely a stubborn, reddish-brown stain. Tilex  
15 Mildew Remover required 30 minutes to remove moderate stains from each of two  
16 undershirts. The third undershirt had a heavy, reddish-brown stain which was much more  
17 stubborn, similar to that treated by Scrubbing Bubbles Mildew Stain Remover, requiring  
18 approximately 75 minutes for complete removal. Lysol Mildew Remover required 30  
19 minutes and Scrub Free Mildew Stain Remover required 20 minutes, respectively, for the  
20 removal of moderate stains.

21 In another experiment, the underarm areas of a 65% polyester and 35% cotton  
22 shirt with underarm stains was soaked in a Scrubbing Bubbles Mildew Stain Remover  
23 spray for the arbitrary period of one hour. The stain was removed with no adverse effects  
24 to the garment.

25 Typical of chlorine-releasing products, such as Tilex Mildew Remover, Lysol  
26 Mildew Remover, Scrubbing Bubbles Mildew Stain Remover and Scrub Free Mildew  
27 Stain Remover, are not safe for use with non-colorfast dyes or with silk cloth. A test was  
28 conducted to study the extent of physical damage to a pair of pure silk male under briefs  
29 soaked in Tilex® Mildew Remover. At 3-1/2 hours the briefs were damaged to shreds.  
30

## 31 EXAMPLE II

1     Comparison of Clorox Bleach to a Cleaning Composition Comprising 2.4 wt % Sodium  
2                     Hypochlorite and 1.25% Sodium Hydroxide

3             Two similar patches (approximately 2.5 x 2.5 cm<sup>2</sup>) of 100% cotton fabric were  
4 cut from the crotch of a new panty. The first patch was immersed in a diluted Clorox  
5 Bleach solution. The diluted Clorox Bleach solution contained about 2.4 wt % sodium  
6 hypochlorite. After six hours of soaking, the first patch showed signs of shredding. After  
7 ten hours of soaking, the first patch shredded completely. In comparison, the second  
8 patch was immersed in a solution which contains about 2.4 wt % sodium hypochlorite  
9 and 1.25 wt % sodium hydroxide. After ten hours of soaking, no effect of shredding was  
10 observed.

11            A test similar to those described in EXAMPLE I was conducted for the solution  
12 that contains 2.4 wt % sodium hypochlorite and 1.25 wt % sodium hydroxide. The  
13 solution was placed in an opaque spray container and used in exactly the same manner  
14 for cleaning panties of menstrual fluid stain as the commercial mildew removers were  
15 used in EXAMPLE I. The solution had essentially the same results and effectiveness in  
16 removing menstrual fluid stains, as compared to the commercial mildew removers used  
17 in EXAMPLE I.

18  
19                                     EXAMPLE III

20     The Damage Effects of Hypochlorite Solutions to Cotton Patches and the Reduction  
21                     Thereof

22            Cotton patches which were resistant to hand-tearing were soaked in different  
23 bleach solutions until damages have begun to occur as evidenced by weakening of the  
24 fabric such that it can be torn by hands with moderate forces. For each bleach solution to  
25 be tested, multiple cotton patches were used. Each patch was inserted into a vial  
26 containing the bleaching solution. The patch was removed periodically from the vial to  
27 determine the extent of damage by manually administering a tearing action. T<sub>C</sub>(D) was  
28 the cumulative time of soaking before the patch became hand-tearable.

29            The bleach solutions were modified from Ultra Clorox Bleach which contains  
30 about 6% NaOCl and less than 0.2% NaOH. Additional NaOH in dry form was added to  
31 Ultra Clorox Bleach to increase the concentration of NaOH. As Table 1 shows, Ultra

1 Clorox Bleach damages cotton fabrics in an accumulated time of approximately one hour.  
2 Decreasing the ratio of NaOCl/NaOH progressively increases the accumulated times for  
3 which the bleach solution is cotton-safe. This Example indicates that NaOH, added to  
4 Ultra Clorox Bleach, can abate the damage of cotton fabrics, thereby rendering the bleach  
5 solution cotton-safe.

6  
7 Table 1. Comparison of the Damage Effects of Bleaching Solutions

8

<b>Cleaning Solution</b>	<b>NaOH (weight percentage)</b>	<b>NaOCl/NaOH (weight percentage ratio)</b>	<b>T<sub>c</sub>(D) (hours)</b>
Ultra Clorox Bleach	0-0.2	over 30:1	1
Solution #1	0.4-0.6	12:1	4
Solution #2	1.0-1.2	5.5:1	6
Solution #3	2.0-2.2	3:1	6
Solution #4	3.0-3.2	2:1	9.5
Solution #5	4.0-4.2	1.5:2	9.5
Solution #6	6.0-6.2	1:1	9.5

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10 The foregoing description of the present invention provides illustration and  
11 description, but is not intended to be exhaustive or to limit the invention to the precise  
12 one disclosed. Modifications and variations are possible consistent with the above  
13 teachings or may be acquired from practice of the invention. Thus, it is noted that the  
14 scope of the invention is defined by the claims and their equivalents.